

## SECTION 6 - NOMENCLATURE

- 1)  $a$  = interfacial area of gas bubble  
(ft<sup>2</sup> bubble surface area/ft<sup>3</sup> expanded liquid)
- 2)  $A$  = cross sectional area of reactor (ft<sup>2</sup>)
- 3)  $c$  = flame-sprayed catalyst thickness (ft)
- 4)  $[ \quad ]$  = concentrations (moles/ft<sup>3</sup>)
- 5)  $[CAT]$  = catalyst concentration (lb cat/ft<sup>3</sup>)
- 6)  $C_L$  = liquid molar concentration (moles/ft<sup>3</sup> liquid)
- 7)  $C_{pi}$  = heat capacity of specie  $i$  (Btu/mole °F)
- 8)  $C_{pO}$  = heat capacity of coolant (Btu/mole °F)
- 9)  $C_{pS}$  = heat capacity of solids (Btu/lb°F)
- 10)  $D$  = tube diameter (ft)
- 11)  $\Delta E$  = activation energy
- 12)  $h_0$  = coolant film resistance [Btu/(hr)(ft<sup>2</sup>)(°F)]
- 13)  $h_s$  = reactant side film resistance [Btu/(hr)(ft<sup>2</sup>)(°F)]
- 14)  $H_{Rj}$  = heat of reaction  $j$  (Btu/mole)
- 15)  $k_G$  = gas phase reaction rate constant (ft/sec)
- 16)  $k_H$  = hydrogenation rate constant [(ft<sup>3</sup><sub>Rx</sub>)<sup>2</sup>/lb cat-hr-mole]

- 17)  $k_L$  = liquid phase reaction rate constant (ft/sec)
- 18)  $k_{L,i}$  = liquid film mass transfer coefficient for specie i  
(ft<sup>3</sup> liquid/ft<sup>2</sup> gas-bubble surface area - sec)
- 19)  $k_o$  = forward rate constant for olefin equilibrium [(ft<sup>3</sup><sub>Rx</sub>)/lb cat-hr]
- 20)  $k^o$  = frequency factor
- 21)  $k_p$  = polymerization rate constant [(ft<sup>3</sup><sub>Rx</sub>)<sup>2</sup>/lb cat-hr-mole]
- 22)  $k_T$  = rate constant at temperature,  $T_a$
- 23)  $k_{WG}$  = water-gas shift forward rate constant [(ft<sup>3</sup><sub>Rx</sub>)<sup>2</sup>/lb cat-hr-mole]
- 24)  $k_c$  = catalyst thermal conductivity [Btu/(hr)(ft<sup>2</sup>)(°F)(ft)]
- 25)  $K_e$  = olefin equilibrium constant
- 26)  $K_i$  = vapor-liquid equilibrium constant for specie i.
- 27)  $K_t$  = thermal conductivity of tube wall [Btu/(hr)(ft<sup>2</sup>)(°F)(ft)]
- 28)  $M(CH_2)_nH$  = active catalyst species with alkyl chain length n attached
- 29)  $M_i$  = molar flow rate of specie i (moles/hr)
- 30)  $M_o$  = molar flow rate of coolant (moles/hr)
- 31)  $N_c$  = number of components
- 32)  $NR$  = number of reactions
- 33)  $N_T$  = number of tubes

- 34)  $p$  = fraction of all monomers that have polymerized.
- 35)  $Q_L$  = heat removed (Btu/hr)
- 36)  $r$  = rate of formation (moles/hr-ft<sup>3</sup><sub>Rx</sub>)
- 37)  $r_j$  = rate of reaction  $j$  (moles/hr-ft<sup>3</sup>)
- 38)  $R$  = gas constant
- 39)  $S_{ij}$  = stoichiometric matrix
- 40)  $t$  = tube-wall thickness (ft)
- 41)  $T$  = temperature (°F)
- 42)  $T_a$  = absolute temperature
- 43)  $T_w$  = temperature at the tube wall (°F)
- 44)  $U$  = overall heat transfer coefficient (Btu/hr-°F-ft<sup>2</sup>)
- 45)  $V_G$  = gas molar flow (moles/sec)
- 46)  $W_S$  = weight flow rate of solids (lbs/hr)
- 47)  $x$  = unit length of reactor (ft)
- 48)  $X_n$  = mole fraction of carbon number species  $n$ .
- 49)  $X_{b,i}$  = mole fraction of specie  $i$  in bulk liquid
- 50)  $Y_i$  = mole fraction of specie  $i$  in gas phase

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